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experiment [][] atom smallest unit into which matter can be divided and still retain the characteristic properties
of an element the word derives from the greek atomos indivisible and the atom was believed to be indivisible
until the early 20th century when electrons and the nucleus were discovered [][] atoms consist of three basic
particles protons electrons and neutrons the nucleus center of the atom contains the protons positively charged
and the neutrons no charge the outermost regions of the atom are called electron shells and contain the
electrons negatively charged $\square\square\square$ an atom consists of a nucleus of protons and generally neutrons surrounded

by an electromagnetically bound swarm of electrons the chemical elements are distinguished from each other by the number of protons that are in their atoms for example any atom that contains 11 protons is sodium and any atom that contains 29 protons is copper [][] 2023[]7[]7[] 2 3 the modern view of atomic structure page id learning objectives to know the meaning of isotopes and atomic masses the precise physical nature of atoms finally emerged from a series of elegant experiments carried out between 1895 and 1915 [][] 2023[]1[]20[] 1 2 atomic structure the nucleus page id objective after completing this section you should be able to describe the basic structure of the atom key terms make certain that you can define and use in context the key terms below atomic number atomic weight electron mass number neutron

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