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consider this situation as it relates to the reaction in model 1 $n_2 + 3h_2 \rightarrow 2nh_3$ a calculate the volumes of reactants consumed and the volume of product made b record the ratio of n_2 to h_2 to nh_3 reduce the ratio to the lowest whole numbers possible

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mole ratio answer $3/4$ mol nh_3 $3/3$ mol h_2 $5/2$ mol h_2 or $3/4$ mol nh_3 $3/3$ mol h_2 $5/2$ mol h_2 2 mol nh_3 2 mol nh_3 key questions 1 what was the given in the example above 2 where did the $3/4$ $3/3$ $5/2$ come from 3 double check to see if $5/2$ moles of h_2 is the correct answer show the math below what happened

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how do the reduced ratios in the last column compare to the coefficients in the reaction in model 1 b use mathematical concepts to explain how your answer in part a is possible 6 the ratio obtained from the coefficients in a balanced chemical equation is called the mole ratio a what is the mole ratio for the reaction in model 1 b

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find the lowest whole number ratios among the moles of elements to do this start by dividing the smallest number of moles into each of the numbers of moles of elements i.e. set the smallest number to 1 this may yield integers or it may yield decimal results that correspond closely to rational fractions

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reduce all ratios to the lowest whole numbers possible 8 circle the phrase below that completes the sentence when two samples contain the same number of atoms the masses of the the ratio of the sample the masses are unrelated samples will be equal masses will be equal to the ratio of the atom's masses 9

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reaction efficiency stoichiometry helps chemists determine the optimal ratio of reactants in order to maximize the yield of a desired product by understanding the stoichiometric relationship chemists can adjust the amount of reactants to ensure the highest possible efficiency of a reaction

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chemists defined a new unit of mass for this reason atomic mass unit amu one twelfth of the mass of a single carbon 12 atom atomic mass often called atomic weight is listed on the periodic table in amu 1 determine the mass of each of the following atoms in amu a c 13 b h 1 c o 17 2

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using the mole ratio pogil answer key students can verify their answers and determine if they have grasped the concept of mole ratios correctly this key provides step by step solutions and explanations for each question allowing students to identify any misconceptions or errors in their reasoning

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