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interpreting a solubility curve this is simply a matter of being able to read the graph correctly and being able to relate what the graph is telling you to the question you are being asked this is best learnt by example 1 what is the lowest temperature at which we can dissolve 20g of the solute in 100g of water

interpreting solubility curves why solubility is a measure of the amount of solute that will dissolve in a given amount of solvent usually water a solubility curve shows how much solute dissolves in a given volume of a solvent at a given temperature how much sugar dissolves in a cup of hot coffee how much salt can dissolve in cold water test pages what is solubility solubility and temperature solubility curves a solubility curve is a graph of solubility measured in g 100 g water against temperature in c solubility in this video i explain how to interpret solubility curves learn the difference between unsaturated saturated and supersaturated solutions along with un interpreting solubility curves

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between grams of solute that can be dissolved in 100 grams of water as the interpreting solubility curves why solubility is a measure of the amount of solute that will dissolve in a given amount of solvent usually water a solubility curve shows how much solute dissolves in a given volume of a solvent at a given temperature how much sugar dissolves in a cup of hot coffee how much salt can dissolve in cold water solubility curves are used to show how the solubility of a substance changes with temperature to read the graph find the line for the substance the amount that dissolves at a given temperature is on the y axis how much KNO_3 dissolves in 100g H_2O at 50°C find the line red find the temperature and follow up to the line green interpreting solubility curves 2.1k 321k views 11 years ago clear simple solubility curves saturated unsaturated supersaturated solutions this video explains how to interpret solubility curves and read to plot a solubility curve the maximum mass of solute that can be dissolved in 100 g of water before a saturated solution is formed is determined at a series of different temperatures solubility curve for three salts while the solubility of most salts increases with temperature sodium chloride or common salt hardly changes at all in this chapter you will learn about the special characteristics of solutions how solutions are characterized and some of their properties skills to develop define these terms solution solute and solvent distinguish solutions mixtures and colloids describe various types of solutions 1.06 triple only understand how to plot and interpret

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solubility curves the solubility of solids changes as temperature changes this can be plotted on a solubility curve the salts shown on this graph are typical the solubility increases as temperature increases starting with volume 66 the series has been published by the journal of physical and chemical reference data in cooperation with the us national institute of standards and technology and is now formally known as the iupac nist solubility data series

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salts increases with temperature sodium chloride or common salt hardly changes at all

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