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mole the number of representative particles in a mole of a substance is 6.02×10^{23} to find the mass of a mole of a compound scientists add together the of the atoms making up the compound atomic mass

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determine the number of moles or grams in the given substance total the atomic masses for all the elements of any compound use the conversion process to get to the desired units

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a 0.5 mol of carbon dioxide CO_2 b 2 mol of oxygen O_2 calculate the number of moles in a 1.505 10²³ atoms of sodium b 1.806 10²⁴ atoms of copper

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the atomic mass of an element expressed in grams is the of a of the element molar mass the mass of a mole of an element is its grams masses to calculate the molar mass of a compound find the number of of each element in one mole of the compound

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moles and molar mass using the information in the table calculate the number of moles in a 2.03 kg sample of citric acid c a 6 h a 8 o a 7 write your answer using three significant figures

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in a similar way chemists use the relationships between the mole and quantities such as mass volume and number of particles to solve problems in chemistry the mole mass relationship how do you convert the mass of a substance to the number of moles of the substance

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introduction the mole is one of the most important concepts in chemistry as it allows for quantitative calculation of amounts of substances that may take part in chemical reactions various formulas are used to calculate the amount of substance in mol m n

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section 10 4 empirical and molecular formulas the percent composition of a compound is the percent by mass of each of the elements in a compound divide the mass of each element in the sample by the mass of the sample then multiply each quotient by 100 c

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the symbol for the unit mole is mol one mole of a substance contains the same number of the stated particles atoms molecules or ions as one mole of any other substance the number of atoms molecules or ions in a mole 1 mol of a given substance is the avogadro constant

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