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chemists seek to control reaction rates so that a desired result is obtained the goal of chemists may be to speed it up so that it proceeds at a fast enough rate to be used in industrial settings the goal of chemists may be to prevent a reaction from going too fast and becoming dangerous equilibrium is indicated when the concentrations no longer change with time the rate of disappearance of N_2O_4 decreases with time as the concentration of N_2O_4 decreases at the same time the rate of formation of NO_2 also decreases with time equilibrium occurs when these two rates are equal this chapter extends the equilibrium discussion of other chapters by addressing some additional reaction classes including precipitation and systems involving coupled equilibrium reactions chapter 15 chemical equilibrium common student misconceptions many students need to see how the numerical problems in this chapter are solved students confuse the arrows used for resonance and equilibrium students often have problems distinguishing between K and Q study with quizlet and memorize flashcards containing terms like aqueous solution solvent solute and more 19 2 coordination chemistry of transition metals 19 3 spectroscopic and magnetic properties of coordination compounds key terms summary exercises calculate change in concentration that occurs as system reaches equilibrium use stoichiometry to determine change in concentration of unknown species from initial concentrations and changes in concentrations calculate equilibrium concentrations chapter 15 1 the concept of chemical equilibrium chapter 15 2 the equilibrium constant chapter 15 3 solving equilibrium problems chapter 15 4 non equilibrium conditions chapter 15 5 factors that affect equilibrium chapter 15 6 controlling the products of reactions chapter 15 7 essential skills chapter 15 8 end of chapter materials ap chemistry chapter 15 equilibrium 1 chapter 15 chemical equilibrium common student misconceptions many students need to see how the numerical problems in this chapter are solved students confuse the arrows used for resonance and equilibrium chapter 15 the properties of solutions page id 342531 15 10 finding the molecular weight of an unknown using colligative properties video 15 11 1 finding the vapor pressure of a solution nonionic nonvolatile solute video 15 11 2 finding the vapor pressure of a solution ionic nonvolatile solute video study with quizlet and memorize flashcards containing terms like solutions solvent solute and more why this chapter 12 1 mass spectrometry of small molecules magnetic sector instruments 12 2 interpreting mass spectra 12 3 mass spectrometry of some common functional groups 12 4 mass spectrometry in biological chemistry time of flight tof instruments 12 5 spectroscopy and the electromagnetic spectrum 12 6 infrared spectroscopy study with quizlet and memorize flashcards containing terms like water is a molecule water has a shape the oxygen atoms of water form a bond with each hydrogen atom and more chapter 15 summary 15 1 salts salts are a chemical compound formed when ions form ionic bonds in these reactions one atom gives up one or more electrons and thus becomes positively charged whereas the other accepts one or more electrons and becomes negatively charged overall the ionic compound has no net charge effect

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definition the acid and base combine to form water neutralizing each
other in the process this video gives you the chemistry of what is
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engineering standpoint 1 determine the molarity of a solution of cucl2
when 540 g cucl2 is dissolved to form a solution with a final volume
of 2 000 ml 2 consider two 1 00 l solutions of 2 00 m cucl2 the first
solution is labeled solution a half 500 ml of the second solution is
poured out into another container this is solution b study with
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of a solution that contains 3 9 10 5 m h at 25 c calculate the poh of
a solution that contains 2 4 10 5 m h at 25 c find the percent
ionization of a 0 337 m hf solution the ka for hf is 3 5 10 4 and more
1 chemical bond the combining of atoms to form molecules or compounds
2 ionic compound a compound made of oppositely charged ions 3 covalent
compound a chemical compound formed by the sharing of electrons
section summary

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