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there are two different isotopes of bromine atoms under normal conditions elemental bromine consists of  $Br_2$  molecules and the mass of a  $Br_2$  molecule is the sum of the masses of the two atoms in the molecule

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the ways in which electrons are arranged into orbitals around the nuclei of atoms are called match the name of the rule used to find the electron configurations of atoms with the rule itself 2 aufbau principle a when electrons occupy orbitals of equal energy one electron enters each orbital until all the orbitals contain one

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this chapter discusses the structure of the atom outlining the sub atomic particles their charges and masses and how atoms can take various forms within their elements known as isotopes students will learn the standard notation for atoms and how to balance nuclear equations in relation to atomic and mass numbers

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a specific amount of energy related to the movement of electrons in atoms f the particle of an atom with a positive charge g the number of protons in the nucleus of every atom of an element h the particle of an atom that is neutral elements and the periodic table 4 1 review and reinforce

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pearson symbol the pearson symbol or pearson notation is used in crystallography as a means of describing a crystal structure and was originated by w b pearson 1 the symbol is made up of two letters followed by a number for example diamond structure cf 8 rutile structure tp 6

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